More Practice Identifying Redox Reactions & Writing Half Reactions WS

Instructions: Determine which of the following equations represent redox reactions. If it is a redox reaction, determine which element is oxidized and which element is reduced.

1)
$$HNO_3 + HI \rightarrow NO + I_2 + H_2O$$

2) C +
$$H_2SO_4 \rightarrow CO_2 + SO_2 + H_2O$$

3)
$$Pb(NO_3)_2 + HCI \rightarrow PbCl_2 + HNO_3$$

Instructions: Write the TWO half reactions from the following equations. Next, determine which equation was oxidation and which was reduction.

1) NbO₂ + W
$$\rightarrow$$
 Nb + WO₄²-

2)
$$Zn + NO_2^{-1} \rightarrow NH_3 + Zn(OH)_4^{2-}$$

3)
$$Sn^{2+} + IO_3^{1-} \rightarrow Sn^{4+} + I^{-1}$$

4)
$$Mn^{2+} + BiO_3^{-1} \rightarrow Bi^{2+} + MnO_4^{-1}$$